

APEAPCET-2026

(Engineering, Agriculture and Pharmacy Common Entrance Test- 2026)

Conducted by

JNT University Kakinada

on behalf of

Andhra Pradesh State Council for Higher Education (APSCHE)

Dates of Examination: 12-05-2026 to 18-05-2026

(No Sessions on 16-05-2026 & 17-05-2026)

(9.00 A.M. to 12.00 Noon & 2.00 P.M to 5.00 P.M)

INSTRUCTION BOOKLET ENGINEERING

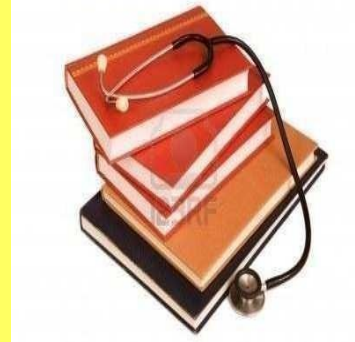
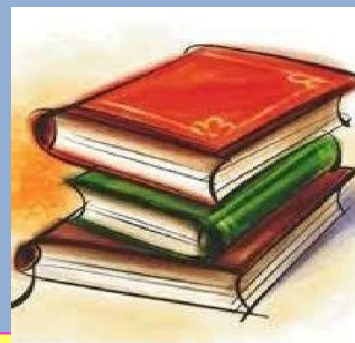
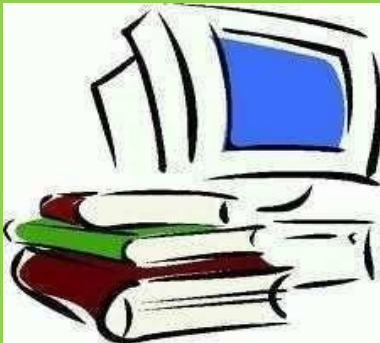
ENGINEERING, AGRICULTURE & PHARMACY COMMON ENTRANCE TEST



JNT University Kakinada
Kakinada



Andhra Pradesh State Council of Higher Education
Tadepalli, Guntur District.



ENGINEERING

ENGINEERING, AGRICULTURE & PHARMACY COMMON ENTRANCE TEST
(being conducted by JNTUK on behalf of APSCHE)

AP EAPCET-2026

FOR ENTRANCE TEST RELATING TO PROFESSIONAL COURSES IN

- a) Engineering, Bio-Technology, B.Tech. (Dairy Technology), B.Tech.(Agr.Engg.), B.Tech. (Food Science and Technology)**
- b) B.Sc. (Ag)/ B.Sc. (Hort)/ B.V.Sc. & A.H/B.F.Sc**
- c) B. Pharmacy, Pharm.D**

Note: Information about the Entrance test is also available on the Website
<https://cets.apsche.ap.gov.in/>

LAST DATES FOR SUBMISSION OF ONLINE APPLICATION	
WITHOUT LATE FEE	07-03-2026
WITH LATE FEE Rs.1000/-	12-03-2026
WITH LATE FEE Rs.2000/-	17-03-2026
WITH LATE FEE Rs.4000/-	21-03-2026
WITH LATE FEE Rs. 10000/-	26-03-2026

Address for Correspondence:

CONVENER, AP EAPCET – 2026,
GROUND FLOOR, ADMINISTRATIVE BUILDING
JAWAHARLAL NEHRU TECHNOLOGICAL UNIVERSITY KAKINADA
KAKINADA – 533003, ANDHRA PRADESH
Contact No: 0884-2359599, 0884-2342499
Email ID: helpdeskapeapcet@apsche.org

AP EAPCET - 2026 (ENGINEERING)

A Common Entrance Test designated as “Engineering, Agriculture & Pharmacy Common Entrance Test” (AP EAPCET – 2026) will be conducted by JNT University Kakinada, Kakinada for the academic year 2026-2027 for admission into the First Year of Professional Courses i.e (i) **Engineering, Bio-Technology, B.Tech. (Dairy Technology), B.Tech. (Agr. Engg.), B.Tech. (Food Science and Technology)**, (ii) **B.Sc. (Ag)/ B.Sc. (Hort)/ B.V.Sc. & A.H/B.F.Sc & (iii) B. Pharmacy, Pharm. D.**

1) PARTICULARS OF AP EAPCET – 2026

- ❖ The Test will be conducted from **12-05-2026 to 18-05-2026 (No Sessions on 16th & 17th May 2026) in two sessions every day i.e. 9.00 A.M. to 12.00 P.M. and 2.00 P.M to 5.00 P.M** through Online mode only.
- ❖ The question paper consists of a total of 160 questions comprising 80 questions in Mathematics, 40 questions in Physics, and 40 questions in Chemistry.
- ❖ All questions are of objective type (multiple choice) only and each question carries one mark. The syllabus in Mathematics, Physics, and Chemistry is furnished in Annexure-I. The model questions are given in Annexure-II.
- ❖ A sample/mock test will be available on the <https://cets.apsche.ap.gov.in> website for practice purposes and to give the candidate a look and feel of the On-Line (Computer Based) Examination.

2) ELIGIBILITY TO APPEAR FOR AP EAPCET – 2026

Candidates satisfying the following requirements shall be eligible to appear for AP EAPCET-2026:

- a. Candidates should be of Indian Nationality or Persons of Indian Origin (PIO) / Overseas Citizen of India (OCI) Card Holders.
- b. Candidates should belong to the State of Andhra Pradesh. The candidates should satisfy Local / Non-Local Status requirements as per the Government of Andhra Pradesh Orders for 2026-27.
- c. Candidate should have obtained at least 45% marks (40% in case of candidate belonging to reserved category) in the subjects specified taken together in the qualifying examination.
- d. For Engineering, B.Pharmacy (M.P.C), Pharm.D, B.Tech. (Dairy Technology), B.Tech. (Agr. Engineering), B.Tech. [Food Science and

Technology (FS & T)], B.Sc. [Agriculture Engg] courses:

- (i) Candidates should have passed Intermediate Examination (10+2 pattern) with Mathematics, Physics, and Chemistry as options or related vocational courses in the fields of Engineering and Technology, conducted by the Board of Intermediate Education, Andhra Pradesh, along with bridge course or courses conducted by it for candidates enrolled from the academic year 2000 onwards, or any other examination recognized as equivalent thereto by the Board of Intermediate Education, Andhra Pradesh. However, the candidates who have appeared for the Final Year Intermediate Examination (10+2 Pattern) and who are awaiting their results may also apply for APEAPCET 2026, but their ranks obtained in APEAPCET 2026 will be valid only if they pass the Intermediate Examination.

OR

- (ii) Candidates should have passed the Diploma examination in Engineering conducted by the State Board of Technical Education and Training, Andhra Pradesh or any other examination recognized as equivalent there to by the State Board of Technical Education and Training, Andhra Pradesh. However, the candidates who have appeared for the Diploma examination in Engineering and who are awaiting their results may also apply for APEAPCET 2026, but their ranks obtained in APEAPCET-2026 will be valid only if they pass the Diploma examination in Engineering.
- (iii) a) In the case of Engineering, Pharmacy courses, candidates should have completed 16 years of age as of 31st December of the year of admission (2026). There is no upper age limit.

b) In the case of B.Tech. (Dairy Technology), B.Tech. (Agr. Engineering), B.Tech. (FS & T) and B.Sc. (Agr. Engg), candidates should have completed 17 years of age as of 31st December of the year of admission (2026) and an upper age limit is 22 years for all the candidates and 25 years in respect of Scheduled Caste and Scheduled Tribe candidates as on 31st December of the year of Admissions (2026).
- e. (i) For Pharm. D course candidates should have passed the Intermediate Examination (10+2 pattern) with Physics, Chemistry, and Mathematics as options conducted by the Board of Intermediate Education, Andhra Pradesh, or any other examination recognized by the Board of Intermediate Education, Andhra Pradesh or equivalent. However, the candidates who have appeared for the Final Year Intermediate Examination (10+2 Pattern) and who are awaiting their results may also apply for AP EAPCET 2026, but their ranks obtained in AP EAPCET 2026 will be valid only if they pass

the Intermediate Examination.

(ii) The candidates should have completed 17 years of age as of 31st December of the year of admission (2026) to the above course.

3) GENERAL INFORMATION / INSTRUCTIONS:

a) The Convener, AP EAPCET – 2026 reserves the right to reject the application of the candidate at any stage, if:

- (I) The Online Application Form is incomplete.
- (II) The candidate fails to satisfy the eligibility conditions.
- (III) Any false or incorrect information is furnished.
- (IV) The Online Application Form is submitted after the due date.

Note: No correspondence will be entertained in this regard.

b) The Convener is not responsible for non-receipt of application by the notified date and time for any reason.

4) MEDIUM OF ENTRANCE TEST:

The question paper contains questions in the “English” and “Telugu” medium only. Candidates, who have studied the qualifying examination in Urdu medium and wish to avail assistance for translating the questions into Urdu, **will be allotted a Test Centre at Kurnool only. In case of ambiguity in the Telugu Question, the Question given in English shall be taken as final.**

5) REGISTRATION FEE:

Payment of Registration Fee for submission of Online Application Form is the first step and the Registration Fee is Rs.800/- for students belonging to the open category (for SC/ST Candidates Rs.700/- and for BC Candidates Rs.750/-) which has to be paid through any kind of Online Payment Modes (Credit Card/ Debit Card/Net Banking/ AP Online etc.,) from any Recognized Bank.

6) SAME CENTRE FOR CANDIDATES APPEARING FOR BOTH ENGINEERING AND AGRICULTURE & PHARMACY: Candidates of E – Category who are eligible and desirous of taking the test in AP- Category, in addition to the test for E - Category should **select the option Both (E & AP Category) together**, during the submission of the Online Application Form, so that same Test Centre can be allotted to them for both the tests. If this instruction is not followed, the candidate may be allotted different Test Centres for E & AP category tests, and Convener, AP EAPCET- 2026 is not responsible for the allotment of different centers.

7) REGIONAL CENTERS

S. No.	New District	Regional Center
1	ANANTAPUR	Anantapur
		Gooty
		Tadipatri
2	SRI SATYASAI	Puttaparthi
3	CHITTOOR	Chittoor
		Kuppam
		Palamner
4	ANNAMAYYA	Madanapalle
		Rayachoty
5	TIRUPATI	Tirupati
		Puttur
6	EAST GODAVARI	Rajahmundry
7	KAKINADA	Kakinada
		Surampalem
8	KONASEEMA	Amalapuram
9	GUNTUR	Guntur
10	BAPATLA	Bapatla
		Chirala
11	PALNADU	Narasaraopeta
12	KRISHNA	Gudlavalleru
		Machilipatnam
13	NTR	Mylavaram
		Tiruvuru
		Vijayawada
14	KURNOOL	Kurnool
		Adoni
		Yemmiganur
15	NANDYAL	Nandyal
16	SRI POTTI SRI RAMULU NELLORE	Kavali
		Vidya Nagar
		Gudur
		Nellore
17	PRAKASAM	Markapuram
		Ongole
18	SRIKAKULAM	Srikakulam
		Tekkali
19	VISHAKAPATNAM	Visakhapatnam
20	ANAKAPALLE	Narsipatnam
21	VIZIANAGARAM	Vizianagaram
		Rajam
22	WEST GODAVARI	Bhimavaram
		Narasapuram
		Tadepalligudem
23	ELURU	Eluru
24	YSR KADAPA	Kadapa
		Rajampet
		Proddatur
25	HYDERABAD	Hyderabad

- Note:** 1. The Convener reserves the right to add or delete some online Test Centers from the list of Regional Centers notified.
2. The Convener reserves the right to allot the candidates to any online Test Centre other than that opted by the candidates.
3. Candidate has to submit not more than one application either for “E” or “AP” or “E&AP” category test. If any candidate submits more than one application for one category, the Convener reserves the right to reject all the applications or accept any one of them.

8. SUBMISSION OF ONLINE APPLICATION FOR AP EAPCET – 2026

Applications should be submitted through Online mode only.

The following information must be kept ready for filling the details during Online submission:

- a. Hall ticket Number of Qualifying Examination
- b. Hall ticket Number of S.S.C. or equivalent
- c. Date of Birth
- d. Caste Certificate in case of SC/ST/BC candidates
- e. Aadhar Number
- f. PH, NCC, Sports etc.
- g. Income Certificate (Up to One Lakh or Up to Two Lakhs or More than Two Lakhs (Rupees))
- h. Ration Card
- i. Study or Residence or relevant certificate for proof of local status (last 12 years)

Online submission:

For Online submission, visit the website <https://cets.apsche.ap.gov.in>. A candidate has to pay Rs.800/- as Registration Fee (for SC/ST Candidates Rs.700/- and for BC Candidates Rs. 750/-) and late fee (if applicable) by opting any kind of Online Payment Modes (Credit Card/ Debit Card/Net Banking/ AP Online / TS Online etc) from any Recognized Bank. After filling out the Online Application Form with the required details, the candidate is required to verify all the details carefully and press Submit button. A filled in Online Application Form will be generated which contains the Registration Number along with filled in details. The candidate is required to take a printout of the Filled-in Online Application Form and it is to be submitted to the Invigilator during the examination **after affixing a recent color photograph duly attested by the Gazetted Officer or Principal of the College where the candidate has studied qualifying examination.** The candidate should use the Registration Number for future correspondence.

Mere appearance and qualifying at AP EAPCET-2026 does not confer any right for admission into professional courses. The candidate has to fulfill the eligibility criteria laid down in the relevant G.O at the time of admission.

9. QUALIFYING MARKS FOR AP EAPCET – 2026

The qualifying percentage of marks for the AP EAPCET-2026 is 25% of the maximum marks considered for the ranking. However, for candidates belonging to Scheduled Caste and Scheduled Tribe, no minimum qualifying mark is prescribed. But their admission will be limited to the extent of seats reserved for such categories (vide G.O.Ms. No. 179, LEN&TE, dated 16.06.1986).

10. AP EAPCET-2026 RESULTS

1. **Evaluation:** Every care will be taken to avoid errors in the evaluation, checking, scrutiny, tabulation, normalization, and ranking.

2. **Ranking:**

- a. Candidates shall be ranked based on the EAPCET normalized marks (75% weightage) and 10+2 (25% weightage) in the order of merit as explained in Annexure-IV and Annexure-V.
- b. The rank obtained in AP EAPCET-2026 is valid for admission to the courses mentioned in the application form for the academic year 2026-2027 only.
- c. The rank card shall be downloaded from the website <https://cets.apsche.ap.gov.in>
- d. Rank obtained with the benefit of relaxation of the minimum qualifying marks at AP EAPCET-2026 by any candidate claiming as SC/ST Category will be canceled in case the claim is found to be invalid at the time of admission to any course of study at any participating University / Institution.

11. The candidates should preserve the Filled-in Online Application Form, the Hall Ticket and the Rank Card to produce them when called for verification.

12. Any malpractice in AP EAPCET-2026 will be dealt with as per rules in force vide G.O.Ms.No: 114, Edn / (IE) Dt: 13th May 1997 for the CET.

13. In any litigation concerning AP EAPCET-2026 Test, Convener is the person to sue and be sued. The Convener (Examination), AP EAPCET – 2026 is not responsible for the allotment of seats at the time of admissions. The Commissioner of Technical Education, Andhra Pradesh is the Convener for the Admissions.

14. Any litigation concerning AP EAPCET-2026 shall be subject to the jurisdiction of the A.P. High Court, Amaravathi only.

15. HALL TICKET

The candidate should download the Hall Ticket from website <https://cets.apsche.ap.gov.in>, and also through Mana Mitra (WhatsApp Governance) of AP.

16. COUNSELLING AND ALLOTMENT OF SEATS

The list of institutions for allotment of candidates with intake in each discipline and category, as per reservations through AP EAPCET – 2026 would be released in the **Information Booklet** for Counseling in due course and the same information would also be released on the website <https://cets.apsche.ap.gov.in>.

IMPORTANT INSTRUCTIONS TO CANDIDATES

1. Material to be brought on the date of examination

Hall Candidate shall bring downloaded Hall Ticket along with any one proof of photo identification viz. Intermediate (10+2) Hall ticket/ Passport/PAN card/ Voter id/ Aadhaar card/ Driving license etc. and a good Ball Point Pen (for rough work, working sheets will be provided by the Test Centre).

2. Other important instructions

- a. Hall ticket issued to the candidate is an important document. Candidates are required to preserve it carefully.
- b. Hall ticket is not transferable. Any tampering of Hall Ticket will automatically lead to the disqualification of the candidate
- c. Candidates shall arrive at the online examination center 2 hours before commencement of the examination. This will enable the candidate to familiarize himself/herself with the online examination process.
- d. Candidate will not be permitted into the Examination Hall once the test commences (i.e. 09:00 AM for FN Session and 02:00 PM for AN Session) and will not be allowed to leave the Examination Hall till the END of the Examination (i.e. 12:00 PM for FN Session and 05:00 PM for AN Session). Candidates will not be allowed into the examination hall even if they are late by one minute.
- e. The candidate does not have the option of choosing a specific date/session to appear for the AP EAPCET- 2026 entrance examination. This information is known to him/her only after downloading Hall Ticket. For any reason, if the candidate fails to appear in the given slot, he/she is treated as absent.
- f. Candidates are required to bring the following to the online examination center:
 - i) Hall Ticket, ii) Filled in Online Application Form, iii) A good Ball Point Pen (for rough work, working sheets will be provided by the Test Centre) and iv) **Attested copy** of Caste certificate (**in case of SC/ST category candidates only**).
- g. Candidates are not allowed to carry any textual material, Calculators, DocuPen, Slide Rules, Log Tables, Electronic Watches with facilities of calculator, printed or written material, bits of papers, mobile phone, pager or any other device, except the Hall Ticket, document as required under point no. 2.(f) inside the Examination Room/Hall. If any candidate is in possession of any of the above items, his/her candidature will be treated as an unfair means and his/her current examination will be canceled & he/she will also be debarred for future examination(s) & the equipment will be seized.

GUIDELINES TO CANDIDATES

1. Please check the Hall ticket carefully for your Name, Date of Birth, Gender, Category, Test Centre Name, Date, and Time of examination.
2. Candidates are advised to reach the venue at least 2 hours before the examination to complete the frisking and registration formalities well before the time. The registration desk will be closed 05 minutes before the examination.
3. The candidate must show, on demand, the Hall Ticket for admission in the examination room/hall. A candidate who does not possess the Hall Ticket issued by the Convener, AP EAPCET-2026, shall not be permitted for the examination under any circumstances by the Centre Superintendent.
4. No candidate, under any circumstances, will be allowed to enter the Examination Centre after the commencement of the examination.
5. A seat indicating the Hall Ticket number will be allocated to each candidate. Candidates should find out and occupy their allotted seats only. Any candidate found to have changed room or the seat on his/her own other than allotted, his/her candidature shall be canceled and no plea would be accepted for it.
6. The candidate should ensure that the question paper is available on the computer in English and Telugu languages only.
7. No Candidate will be allowed to carry any baggage inside the Examination Centre. The Convener, AP EAPCET-2026 will not be responsible for any belongings stolen or lost at the premises.
8. Smoking and eating are strictly prohibited in the examination room.
9. Tea, coffee, cool drinks, or snacks are not allowed to be taken into the examination rooms during examination hours.
10. Approach the Centre Superintendent/Invigilator in the room for any technical assistance, first aid emergency, or any other information during the examination.
11. No candidate, without the special permission of the Centre Superintendent or the Invigilator concerned, will leave his/her seat or Examination Room until the full duration of the Examination. Candidates must follow the instructions strictly as instructed by the Centre Superintendent/Invigilators.
12. For any queries or issues regarding computer-based examination, the candidates may contact helpline numbers which will be available on the <https://cets.apsche.ap.gov.in> website.

INSTRUCTIONS FOR ONLINE (COMPUTER BASED) EXAMINATION

The On-Line (Computer Based) Examination will be conducted as per the schedule.

1. The test will start exactly at the time mentioned in the Hall Ticket and an announcement to this effect will be made by the invigilator.
2. The Entrance test is conducted for a duration of 3 hours and the question paper consists of a total of 160 questions comprising 80 questions in Mathematics, 40 questions in Physics, and 40 questions in Chemistry. All questions are having equal weightage.
3. There is only one correct response for each question out of four responses given.
4. There is no negative marking and No deduction from the total score will be made if no response is indicated for a question.
5. All calculations/writing work is to be done only in the rough sheet provided at the center and on completion of the test, candidates must hand over the rough sheets to the invigilator on duty in the Room/Hall. The candidates shall write their Hall Ticket number on the rough sheets used by them.
6. During the examination time, the invigilator will check the Hall ticket of the candidate to satisfy himself/herself about the identity of each candidate.
7. The candidates are governed by all Rules and Regulations of the Convener, EAPCET-2026 with regard to their conduct in the Examination Hall. All cases of unfair means will be dealt with as per rules.

The following Proforma I, II and III are to be submitted at the time of counseling to claim nativity, community and local status.

PROFORMA – I

**REVISED PROFORMA AS PER G.O.Ms.No.58, SOCIAL WELFARE (J) DEPT.
DATED 12.05.1997 ANDHRA PRADESH GAZETTE EXTRAORDINARY PART-I
Serial No. FORM III**

S.C.

District Code :

S.T.

Emblem

Mandal Code:

B.C.

Village Code:

Certificate No.:

**COMMUNITY, NATIVITY AND DATE OF BIRTH CERTIFICATE
(Integrated Community Certificate)**

1. This is to Certify that Sri / Smt / Kum _____
Son / Daughter of Sri _____ of village /Town
_____Mandal_____District of the state of Andhra Pradesh
belongs to _____Community which was recognized as
SC/ST/BC under
The Constitution (Scheduled Caste) Order, 1950
The Constitution (Scheduled Tribes) Order, 1950
G.O.Ms.No.1793, Education, dated 25.09.1970 as amended from time to time BCs, SCs, STs
list (Modification) Order 1956, SCs and STs (Amendment) Act, 1976.

2. It is to certify that Sri / Smt / Kum _____ is a native of _____
the District of Andhra Pradesh.

3. It is to certify that the date of birth of Sri / Smt / Kum _____ is
Day _____Month _____Year _____(in words
_____)
as per the declaration given by his /her Father / Mother / Guardian and as entered in the
school records where he/she studied

(Seal)

Signature

Date

Name in Capital letters:

Designation:

Explanatory Note:

1) While mentioning the community, the competent Authority must mention the sub-caste (in case of SCs) and Sub-Tribe or Sub- Group (in case of STs) as listed out in the SCs and STs (Amendment) Act, 1976.

PROFORMA – II
RESIDENCE CERTIFICATE IN SUPPORT OF APPLICATION

As per the format issued by the Government of Andhra Pradesh for 2026-27

PROFORMA – III

CERTIFICATES IN SUPPORT OF NON-LOCAL STATUS FOR E - CATEGORY

As per the format issued by the Government of Andhra Pradesh for 2026-27

ANNEXURE - I

AP EAPCET – 2026 SYLLABUS

NOTE:

- ❖ The details of the syllabus in which the exam will be conducted are furnished below for the convenience of students.
- ❖ The syllabus is applicable to students of both the current and previous batches of Intermediate Course, who desire to appear for AP EAPCET-2026.

APEAPCET 2026 - MATHEMATICS SYLLABUS

SUBJECT: MATHEMATICS

ALGEBRA

- a) **Functions:** Types of functions – Definitions - Inverse functions & Theorems - Domain, Range and Inverse.
- b) **Mathematical Induction:** Principles of Mathematical Induction & Theorems – Applications of Mathematical Induction – Problems on divisibility.
- c) **Matrices:** Types of matrices - Scalar multiple of a matrix and multiplication of matrices - Transpose of a matrix – Determinants - properties of determinants - Adjoint and Inverse of a matrix – Consistency and inconsistency of system of simultaneous equations - Rank of a matrix - Solution of simultaneous linear equations.
- d) **Complex Numbers:** Complex number as an ordered pair of real numbers- fundamental operations - Representation of complex numbers in the form $a+ib$ - Modulus and amplitude of complex numbers–Illustrations - Geometrical and Polar Representation of complex numbers in Argand plane-Argand diagram.
- e) **De Moivre's Theorem:** De Moivre's theorem- Integral and Rational indices - n^{th} roots of unity- Geometrical Interpretations–Illustrations.
- f) **Quadratic Expressions:** Quadratic expressions, equations in one variable - Sign of quadratic expressions – Change in signs – Maximum and minimum values - Quadratic Inequations.
- g) **Theory of Equations:** The relation between the roots and coefficients in an equation - Solving an equations when two or more roots of it are connected by certain relation - Equation with real coefficients, occurrence of complex roots in conjugate pairs and its consequences, Transformation of equations- Reciprocal equations.
- h) **Permutations and Combinations:** Fundamental Principle of counting – linear and circular permutations- Permutations of 'n' dissimilar things taken 'r' at a time - Permutations when repetitions are allowed - Circular permutations - Permutations with constraint repetitions - Combinations-definitions, certain theorems.
- i) **Binomial Theorem:** Binomial theorem for positive integral index, Binomial theorem for rational Index - Approximations using Binomial theorem
- j) **Partial fractions:** Partial fractions of $f(x)/g(x)$ when $g(x)$ contains non –repeated linear factors - Partial fractions of $f(x)/g(x)$ where both $f(x)$ and $g(x)$ are polynomials and when $g(x)$ contains repeated and/or non-repeated linear factors - Partial fractions of $f(x)/g(x)$ when $g(x)$ contains irreducible factors, Partial fractions of $f(x)/g(x)$ when $f(x)/g(x)$ is an improper fraction, conversion of $f(x)/g(x)$ in power series of x .

TRIGONOMETRY

- a) **Trigonometric Ratios upto Transformations:** Trigonometric ratios – Variation - Graphs and Periodicity of Trigonometric functions - Trigonometric ratios of Compound angles - Trigonometric ratios of multiple and sub- multiple angles - Sum and Product transformations.
- b) **Trigonometric Equations:** General solutions of Trigonometric Equations – Simple Trigonometric Equations – Solutions.
- c) **Inverse Trigonometric Functions:** To reduce a Trigonometric function into a bijective function – Graphs of Inverse Trigonometric functions – Properties of Inverse Trigonometric functions.
- d) **Hyperbolic Functions:** Definitions of Hyperbolic Functions – Graphs - Definitions of Inverse

Hyperbolic Functions – Graphs - Addition formulae of Hyperbolic Functions.

- e) **Properties of Triangles:** Relation between sides and angles of a Triangle - Sine, Cosine, Tangent and Projection rules- Half angle formulae and areas of a triangle – Incircle and Excircles of a Triangle.

VECTOR ALGEBRA

- a) **Addition of Vectors:** Vectors as a triad of real numbers – Some Basic Concepts - Classification of vectors - Addition of vectors - Scalar multiplication - Angle between two non-zero vectors - Linear combination of vectors - Components of a vector in three dimensions - Vector equations of line and plane including the Cartesian equivalent form of line.
- b) **Product of Vectors:** Scalar or dot product of two vectors - Geometrical Interpretations - orthogonal projections - Properties of dot product - Expression of dot product in (i, j, k) system - Angle between two vectors - Geometrical Vector methods – Vector equations of plane in normal form-Angle between two planes- Vector product of two vectors and properties- Vector product in (i, j, k) system- Vector Areas – Scalar triple product – Vector equation of a plane – different forms, skew lines, shortest distance – plane, condition for coplanarity etc. – vector triple product – results.

MEASURES OF DISPERSION AND PROBABILITY

- a) **Measures of Dispersion** - Range - Mean deviation - Variance and standard deviation of ungrouped/grouped data, coefficient of variation and analysis of frequency distributions with equal means and unequal means but different variances.
- b) **Probability:** Random experiments and events - Classical definition of probability, Axiomatic approach and addition theorem of probability - Independent and dependent events - conditional probability- multiplication theorem and Baye's theorem.
- c) **Random Variables and Probability Distributions:** Random Variables - Theoretical discrete distributions - mean and variance of discrete random variable – Binomial and Poisson Distributions.

COORDINATE GEOMETRY

- a) **Locus:** Definition of locus –Illustrations-To find equations of locus-Problems connected to it.
- b) **Transformation of Axes:** Transformation of Axes – Rules, derivations and illustrations – translation of axes – Rotation of Axes – Derivations – Illustrations.
- c) **The Straight Line:** Revision of fundamental results - Straight line - Normal form – Illustrations - Straight line - Symmetric form - Straight line - Reduction into various forms - Intersection of two Straight Lines - Family of straight lines - Concurrent lines - Condition for Concurrent lines - Angle between two lines - Length of perpendicular from a point to a Line - Distance between two parallel lines - Concurrent lines - properties related to a triangle.
- d) **Pair of Straight lines:** Equations of pair of lines passing through origin - angle between a pair of lines - Condition for perpendicular and coincident lines, bisectors of angles - Pair of bisectors of angles - Pair of lines - second degree general equation - Conditions for parallel lines - distance between them, Point of intersection of pair of lines - Homogenising a second degree equation with a first degree equation in x and y.
- e) **Circle :** Equation of circle -standard form-centre and radius - Equation of a circle with a given line segment as diameter & equation of circle through three non collinear points - parametric equations of a circle - Position of a point in the plane of a circle – power of a point-definition of tangent-length of tangent - Position of a straight line in the plane of a circle-conditions for a line to be tangent – chord joining two points on a circle – equation of the tangent at a point on the circle- point of contact-equation of normal-Chord of contact-pole and polar-conjugate points and conjugate lines- equation of chord with given middle point, Relative positions of two

circles- circles touching each other externally, internally common tangents –centers of similitude- equation of pair of tangents from an external point.

- f) **System of circles:** Angle between two intersecting circles –condition for orthogonality - Radical axis of two circles- properties- Common chord and common tangent of two circles – radical centre - Intersection of a line and a Circle.
- g) **Parabola:** Conic sections – Conic – Parabola- equation of parabola in standard form – Definitions of Chord, Focal Chord, Double Ordinate, latus rectum - different forms of parabola- parametric equations, parabola and point in the plane - Equations of tangent and normal at a point on the parabola (Cartesian and Parametric)- conditions for straight line to be a tangent.
- h) **Ellipse:** Equation of ellipse in standard form- Parametric equations, various forms of ellipse - Equation of tangent and normal at a point on the ellipse (Cartesian and parametric)- condition for a straight line to be a tangent – auxiliary circles, eccentric angles, director circle.
- i) **Hyperbola:** Equation of hyperbola in standard form- rectangular hyperbola – various forms of hyperbola - auxiliary circle, Parametric equations - Equations of tangent and normal at a point on the hyperbola (Cartesian and parametric) - conditions for a straight line to be tangent- Asymptotes – Director circle of a hyperbola.
- j) **Three Dimensional Coordinates:** Coordinates - Section formulae - Centroid of a triangle and tetrahedron.
- k) **Direction Cosines and Direction Ratios:** Direction Cosines –Direction Ratios.
- l) **Plane:** Cartesian equation of a Plane –Simple Illustrations – angle between two planes.

CALCULUS

- a) **Limits and Continuity:** Intervals and neighbourhoods – Limits - Standard Limits–Continuity.
- b) **Differentiation:** Derivative of a function - Elementary Properties - Trigonometric, Inverse Trigonometric, Hyperbolic, Inverse Hyperbolic Functions – Derivatives - Methods of Differentiation – Second Order Derivatives.
- c) **Applications of Derivatives:** Errors & Approximations - Geometrical Interpretation of a derivative - Equations of tangents and normal to a curve – Lengths of Tangent, Normal, Subtangent and subnormal - Angles between two curves and condition for orthogonality of curves – Derivative as a rate of change – Rolle’s theorem and Lagrange’s Mean value theorem - Increasing and decreasing functions - Maxima and Minima.
- d) **Integration:** Integration as the inverse process of differentiation- Standard forms -properties of integrals - Method of substitution- integration of Algebraic, exponential, logarithmic, trigonometric and inverse trigonometric functions - Integration by parts – Integration by the method of substitution – Integration of algebraic and trigonometric functions – Integration by parts – Integration of exponential, logarithmic and inverse trigonometric functions – Integration - Partial fractions method – Reduction formulae.
- e) **Definite Integrals:** Definite Integral as the limit of sum, Interpretation of Definite Integral as an area. Fundamental theorem of Integral Calculus. Properties, Reduction formulae, Application of Definite integral to areas.
- f) **Differential equations:** Formation of differential equation-Degree and order of an ordinary differential equation – Solution of a differential equation - Solving differential equation by
 - i) Variables separable method, ii) Homogeneous differential equation, iii) Non Homogeneous differential equation iv) Linear differential equations

APEAPCET 2026 - PHYSICS SYLLABUS

SUBJECT: PHYSICS

1. PHYSICAL WORLD: What is physics? Scope and excitement of physics. Physics, technology and society, Fundamental forces in nature, Nature of physical laws

2. UNITS AND MEASUREMENTS: The international system of units, Measurement of Length, Measurement of Large Distances, Estimation of Very Small Distances, Size of a Molecule, Range of Lengths, Measurement of Mass, Range of Masses, Measurement of time, Accuracy, precision of instruments and errors in measurement, Systematic errors, random errors, least count error, Absolute Error, Relative Error and Percentage Error, Combination of Errors, Significant figures, Rules for Arithmetic Operations with Significant Figures, Rounding off the Uncertain Digits, Rules for Determining the Uncertainty in the Results of Arithmetic Calculations, Dimensions of Physical Quantities, Dimensional Formulae and dimensional equations, Dimensional Analysis and its Applications, Checking the Dimensional Consistency of Equations, Deducing Relation among the Physical Quantities.

3. MOTION IN A STRAIGHT LINE: Position, path length and displacement, average velocity and average speed, instantaneous velocity and speed, acceleration, kinematic equations for uniformly accelerated motion, relative velocity.

4. MOTION IN A PLANE: Scalars and vectors, position and displacement vectors, equality of vectors, multiplication of vectors by real numbers, addition and subtraction of vectors - graphical method, resolution of vectors, vector addition - analytical method, motion in a plane, position vector and displacement, velocity, acceleration, motion in a plane with constant acceleration, relative velocity in two dimensions, projectile motion, equation of path of a projectile, time of maximum height, maximum height of a projectile, horizontal range of projectile, uniform circular motion.

5. LAWS OF MOTION: Aristotle's fallacy, The Law of inertia, Newton's first law of motion, Newton's second law of motion- momentum, impulse, Newton's third law of motion, conservation of momentum, Equilibrium of a particle, Common forces in mechanics, friction, types of friction, static, kinetic and rolling frictions, Circular motion, Motion of a car on a level road, Motion of a car on a banked road, solving problems in mechanics.

6. WORK, ENERGY AND POWER: The Scalar Product, Notions of work and kinetic energy, The work-energy theorem, Work, Kinetic energy, Work done by a variable force, The work-energy theorem for a variable force, The concept of Potential Energy, The conservation of Mechanical Energy, The Potential Energy of a spring, Various forms of energy, Heat, Chemical Energy, Electrical Energy, The Equivalence of Mass and Energy, Nuclear Energy, The Principle of Conservation of Energy, Power, Collisions, Elastic and Inelastic Collisions, Collisions in one dimension, Coefficient of Restitution and its determination, Collisions in Two Dimensions.

7. SYSTEM OF PARTICLES AND ROTATIONAL MOTION: Rigid body motion, Centre of mass, Centre of Gravity, Motion of centre of mass, Linear momentum of a system of particles, Vector product of two vectors, Angular velocity and its relation with linear velocity, Angular acceleration, Kinematics of rotational motion about a fixed axis, Moment of force (Torque), Angular momentum of particle, Torque and angular momentum for a system of a particles - conservation of angular momentum, Equilibrium of a rigid body, Principle of moments, Moment of inertia, Dynamics of rotational motion about a fixed axis, Angular momentum in case of rotation about a fixed axis - conservation of angular momentum, Rolling motion, Kinetic Energy of Rolling Motion.

8. OSCILLATIONS: Periodic and oscillatory motions, Period and frequency, Displacement, Simple harmonic motion (S.H.M.), Simple harmonic motion and uniform circular motion, Velocity and acceleration in simple harmonic motion, Force law for Simple harmonic Motion, Energy in simple harmonic motion, some systems executing Simple Harmonic Motion, Oscillations due to a spring, The Simple Pendulum, damped simple harmonic motion, Forced oscillations and resonance.

9. GRAVITATION: Kepler's laws Law of Orbits, Law of Areas, Law of Periods, Universal law of gravitation, central forces, the gravitational constant, Acceleration due to gravity of the earth, Acceleration due to gravity below and above the surface of earth, Gravitational potential energy, Escape speed, Earth satellites, Energy of an orbiting satellite, Geostationary and polar satellites, Weightlessness.

10. MECHANICAL PROPERTIES OF SOLIDS: Elastic behavior of solids, Stress and strain, Hooke's law, Stress-strain curve, Elastic moduli, Young's Modulus, Determination of Young's Modulus of the Material of a Wire, Shear Modulus, Bulk Modulus, Poisson's ratio, Elastic potential energy in a stretched wire, Applications of elastic behavior of materials.

11. MECHANICAL PROPERTIES OF FLUIDS: Pressure, Pascal's Law, Variation of Pressure with Depth, Atmospheric Pressure and Gauge Pressure, Hydraulic Machines, Archimedes' Principle, Streamline flow, Bernoulli's principle, Speed of Efflux, Torricelli's Law, Venturi- meter, Blood Flow and Heart Attack, Dynamic Lift, Viscosity, Variation of Viscosity of fluids with temperature, Stokes' Law, Reynolds number, Critical Velocity, Surface tension and Surface Energy, Angle of Contact, Drops and Bubbles, Capillary Rise, Detergents and Surface Tension.

12. THERMAL PROPERTIES OF MATTER: Temperature and heat, Measurement of temperature, Ideal-gas equation and absolute temperature, Thermal expansion, Specific heat capacity, Calorimetry, Change of state, Triple Point, Regelation, Latent Heat, Heat transfer – Conduction, convection and radiation, Black body Radiation, Greenhouse Effect, Newton's law of cooling and its experimental verification.

13. THERMODYNAMICS: Thermal equilibrium, Zeroth law of thermodynamics, Heat, Internal Energy and work, First law of thermodynamics, Specific heat capacity, Specific heat capacity of water, Thermodynamic state variables and equation of State, Thermodynamic processes, Quasi-static process, Isothermal Process, Adiabatic Process, Isochoric Process, Isobaric process, Cyclic process, Heat engines , Refrigerators and heat pumps, Second law of thermodynamics, Reversible and irreversible processes, Carnot engine, Carnot's theorem.

14. KINETIC THEORY: Molecular nature of matter, Behaviour of gases, Boyle's Law, Charles' Law, Kinetic theory of an ideal gas, Pressure of an Ideal Gas, Kinetic interpretation of temperature, Law of equipartition of energy, Specific heat capacity, Monatomic Gases, Diatomic Gases, Polyatomic Gases, Specific Heat Capacity of Solids, Specific Heat Capacity of Water, Mean free path.

15. WAVES: Transverse and longitudinal waves, wave displacement relation in a progressive wave, amplitude and phase, wavelength and angular wave number, period, angular frequency and frequency, the speed of a travelling wave, speed of a transverse wave on stretched string, speed of a longitudinal wave (speed of sound), the principle of superposition of waves, reflection of waves, standing waves and normal modes, beats, Doppler effect – source moving & Observer stationary, observer moving and source stationary, both observer and source are moving, applications of Doppler effect.

16. RAY OPTICS AND OPTICAL INSTRUMENTS: Reflection of Light by Spherical Mirrors, Sign convention, Focal length of spherical mirror, Mirror equation, refraction, total internal reflection, total internal reflection in nature and its technological applications, refraction at spherical surfaces and by lenses, power of a lens, combination of thin lenses in contact, refraction through a prism, dispersion by a prism, natural phenomena due to sunlight – Rainbow, Scattering of light, optical instruments, the eye, the simple and compound microscopes, refracting telescope and Cassegrain reflecting telescope.

17. WAVE OPTICS: Huygens principle, refraction and reflection of plane waves using Huygens principle: Refraction of a plane wave, Refraction in a rarer medium (at the denser medium boundary), reflection of a plane wave by a plane surface, the Doppler effect, coherent and incoherent addition of waves, interference of light waves and Young's experiment, Diffraction, Single slit, seeing the single slit diffraction pattern resolving power of optical instruments, Determination of resolving power of eye, the validity of ray optics, Polarization: Polarization by scattering, Polarisation by reflection.

18. ELECTRIC CHARGES AND FIELDS: Electric charge, conductors and insulators, charging by induction, basic properties of electric charges, additivity of charges, conservation of charge, quantization of charge, Coulomb's law, forces between multiple charges, electric field, electric field due to a system of charges, physical significance of electric field, electric field lines, electric flux, electric dipole, the field of an electric dipole for points on the axial line and on the equatorial plane, physical significance of dipoles, dipole in a uniform external field, continuous charge distribution, Gauss's law, Applications of Gauss's Law: Field due to an infinitely long straight uniformly charged wire, field due to uniformly charged infinite plane sheet, field due to uniformly charged thin spherical shell.

19. ELECTROSTATIC POTENTIAL AND CAPACITANCE: Electrostatic potential, potential due to a point charge, potential due to an electric dipole, potential due to a system of charges, equipotential surfaces, relation between field and potential, potential energy of a system of charges, potential energy in an external field, potential energy of a single charge, potential energy of a system of two charges in an external field, potential energy of a dipole in an external field, electrostatics of conductors, dielectrics and polarisation, electric displacement, capacitors and capacitance, the parallel plate capacitor, effect of dielectric on capacitance, combination of capacitors, capacitors in series, capacitors in parallel, energy stored in a capacitor, Van de Graaff generator.

20. CURRENT ELECTRICITY: Electric current, electric current in conductors, Ohm's law, drift of electrons and the origin of resistivity, mobility, limitations of Ohm's law, resistivity of various materials, colour code of resistors, Temperature dependence of resistivity, electrical energy, power. Combination of Resistors, Series and Parallel, Cells, EMF, internal resistance, cells in series and in parallel, Kirchhoff's rules, Wheatstone Bridge, Meter Bridge, Potentiometer.

21. MOVING CHARGES AND MAGNETISM: Magnetic force, sources and fields, magnetic field, Lorentz force, magnetic force on a current carrying conductor, motion in a magnetic field, helical motion of charged particles, motion in combined electric and magnetic fields, , velocity selector, cyclotron, magnetic field due to a current element, Biot – Savart's law, Magnetic field on the axis of a circular current loop, Ampere's circuital law, the solenoid and the toroid, force between two parallel current carrying conductors, the ampere (UNIT), torque on current loop, magnetic dipole, torque on a rectangular current loop in a uniform magnetic field, circular current loop as a magnetic dipole, the magnetic dipole moment of a revolving electron, the Moving Coil Galvanometer; conversion into ammeter and voltmeter.

22. MAGNETISM AND MATTER: The bar magnet, the magnetic field lines, bar magnet as an equivalent solenoid, The dipole in a uniform magnetic field, the electrostatic analog, Magnetism and Gauss's Law, The Earth's magnetism, magnetic declination and dip, magnetization and

magnetic intensity, magnetic properties of materials – Diamagnetism, Paramagnetism and Ferromagnetism, permanent magnets and electromagnets.

23. ELECTROMAGNETIC INDUCTION: The experiments of Faraday and Henry, magnetic flux, Faraday's Law of induction, Lenz's law and conservation of energy, motional electromotive force, energy consideration - a quantitative study, Eddy currents, inductance, mutual inductance, self-inductance, AC generator.

24. ALTERNATING CURRENT: AC voltage applied to a resistor, representation of AC current and voltage by rotating vectors - Phasors, AC voltage applied to an inductor, AC voltage applied to a capacitor, AC voltage applied to a series LCR circuit, Phasor – diagram solution, analytical solution, resonance, sharpness of resonance, Power in AC circuit: The power factor, Wattless current LC oscillations, transformers.

25. ELECTROMAGNETIC WAVES: Displacement Current, Maxwell's equations, electromagnetic waves, sources of electromagnetic waves, nature of electromagnetic waves, electromagnetic spectrum: radio waves, microwaves, infrared waves, visible rays, ultraviolet rays, X-rays, gamma rays.

26. DUAL NATURE OF RADIATION AND MATTER: Electron emission, Photoelectric Effect, Hertz's observations, Hallwachs and Lenard's observations, experimental study of photoelectric effect, effect of intensity of light on photocurrent, effect of potential on photoelectric current, effect of frequency of incident radiation on stopping potential, Photoelectric effect and Wave theory of Light, Einstein's Photoelectric equation, Energy Quantum of Radiation, particle nature of light, the photon, wave nature of matter, photocell, Davisson and Germer Experiment

27. ATOMS: Alpha particle scattering and Rutherford's nuclear model of atom, alpha particle trajectory, electron orbits, atomic spectra, spectral series, Bohr model of the hydrogen atom, energy levels, Franck – Hertz experiment, the line spectra of the hydrogen atom, deBroglie's explanation of Bohr's second postulate of quantization.

28. NUCLEI: Atomic masses and composition of nucleus, discovery of neutron, size of the nucleus, Mass - Energy, Nuclear Binding Energy, The Binding energy per Nucleon and its variation with Mass Number, Nuclear Force, Radioactivity - Law of radioactive decay, half-life and mean life of a Radioactive material, Alpha decay, Beta decay and Gamma decay, Nuclear Energy, Fission, Nuclear reactor, nuclear fusion, energy generation in stars, controlled thermonuclear fusion.

29. SEMICONDUCTOR ELECTRONICS: MATERIALS, DEVICES AND SIMPLE CIRCUITS: Classification of metals, conductors, and semiconductors on the basis of conductivity and energy bands, Band theory of solids, Intrinsic semiconductor, Extrinsic semiconductor, p-type semiconductor, n-type semiconductor, p-n junction, forward bias, reverse bias, Semiconductor diode, Application of junction diode as a rectifier, Zener Diode, Zener Diode as a voltage regulator, Optoelectronic junction devices, Photodiode, light emitting diode, solar cell. Junction transistor, structure and action, Basic transistor circuit configurations and transistor characteristics, transistor as a switch and as an amplifier (CE – Configuration), Feedback amplifier and transistor oscillator, Digital Electronics and Logic gates, NOT, OR, AND, NAND and NOR Gates, Integrated circuits.

30. COMMUNICATION SYSTEMS: Elements of a Communication system, basic terminology used in electronic communication systems, bandwidth of signals, bandwidth of transmission medium, propagation of electromagnetic waves, ground waves, sky waves, space wave, modulation and its necessity, size of the antenna or aerial, effective power radiated by an antenna, mixing up of signals from different transmitters, amplitude modulation, production of amplitude modulated wave, detection of amplitude modulated wave.

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APEAPCET 2026 - CHEMISTRY SYLLABUS

SUBJECT: CHEMISTRY

Unit-1: ATOMIC STRUCTURE: Sub Atomic particles, Atomic models, Developments to the Bohr's model of atom; Wave nature of electromagnetic radiation; Particle nature of electromagnetic radiation, Planck's quantum theory; Evidence for the quantized electronic Energy levels : Atomic spectra, Bohr's model for Hydrogen atom; Explanation of line spectrum of hydrogen; Limitations of Bohr's model; Quantum mechanical considerations of sub atomic particles; Dual behaviour of matter; Heisenberg's uncertainty principle; Quantum mechanical model of an atom. Important features of Quantum mechanical model of atom; Orbitals and quantum numbers; Shapes of atomic orbitals; Energies of orbitals; Filling of orbitals in atoms. Aufbau Principle, Pauli's exclusion Principle and Hund's rule of maximum multiplicity; Electronic configurations of atoms; Stability of half-filled and completely filled orbitals.

Unit-2: CLASSIFICATION OF ELEMENTS AND PERIODICITY IN PROPERTIES: Need to classify elements - Genesis of periodic classification, Modern periodic law and present form of the periodic table; Nomenclature of elements with atomic number greater than 100; Electronic configuration of elements and the periodic table; Electronic configuration and types of elements s, p, d and f blocks; Trends in physical properties: (a) Atomic radius, (b) Ionic radius (c) Variation of size in inner transition elements, (d) Ionization enthalpy, (e) Electron gain enthalpy, (f) Electro negativity; Periodic trends in chemical properties: (a) Periodicity of Valence or Oxidation states, (b) Anomalous properties of second period elements –diagonal relationship; Periodic trends and chemical reactivity.

Unit-3: CHEMICAL BONDING AND MOLECULAR STRUCTURE: Kossel - Lewis approach to chemical bonding, Octet rule, Lewis representation of simple molecules, formal charges, limitations of octet rule; Ionic or electrovalent bond - Factors favourable for the formation of ionic compounds- Crystal structure of sodium chloride, Lattice Enthalpy: General properties of ionic compounds; Bond Parameters - bond length, bond angle, and bond enthalpy, bond order, resonance- Polarity of bonds dipole moment-Fajan rules; Valence Shell Electron Pair Repulsion (VSEPR) theory; Predicting the geometry of simple molecules; Valence bond theory-Orbital overlap concept- Directional properties of bonds-overlapping of atomic orbitals- types of overlapping and nature of covalent bonds-strength of sigma and pi bonds-Factors favouring the formation of covalent bonds; Hybridisation- different types of hybridization involving s, p and d orbitals- shapes of simple covalent molecules; Coordinate bond - definition with examples; Molecular orbital theory - Formation of molecular orbitals, Linear combination of atomic orbitals(LCAO)-conditions for combination of atomic orbitals-, Types of Molecular orbitals, Energy level diagrams for molecular orbitals -, Electronic Configuration and Molecular Behaviour, Bonding in some homo nuclear diatomic molecules- H_2 , He_2 , Li_2 , B_2 , C_2 , N_2 and O_2 ; Hydrogen bonding-cause of formation of hydrogen bond - Types of hydrogen bonds-inter and intra molecular-General properties of hydrogen bonds.

Unit-4: STATES OF MATTER: GASES AND LIQUIDS: Intermolecular forces; Thermal Energy; Intermolecular forces Vs Thermal interactions; The Gaseous State; The Gas Laws; Ideal gas equation; Graham's law of diffusion - Dalton's Law of partial pressures; Kinetic molecular theory of gases; Kinetic gas equation of an ideal gas (No derivation) deduction of gas laws from Kinetic gas equation; Distribution of molecular speeds, Kinetic Energy, Behaviour of real gases - Deviation from Ideal gas behaviour - Compressibility factor Vs Pressure diagrams of real gases; Liquification of gases, Liquid state, Vapour Pressure, Surface tension, Viscosity (No mathematical part).

Unit-5: STOICHIOMETRY: Significant figures, Laws of Chemical Combinations - Law of

Conservation of Mass, Law of Definite Proportions, Law of Multiple Proportions, Atomic and molecular masses- mole concept and molar mass. Concept of equivalent weight; Percentage composition of compounds and calculations of empirical and molecular formulae of compounds; Stoichiometry and stoichiometric calculations- limiting reagent; Methods of Expressing concentrations of solutions- mass percent, mole fraction, molarity, molality and normality; Redox reactions-classical idea of redox reactions, oxidation and reduction reactions- redox reactions in terms of electron transfer; Oxidation number concept; Types of Redox reactions- combination, decomposition, displacement and disproportionation reactions; Balancing of redox reactions- oxidation number method Half reaction (ion-electron)method; Redox reactions in titrimetry.

Unit-6: THERMODYNAMICS: Thermodynamic Terms; The system and the surroundings; Types of systems and surroundings; The state of the system; The Internal Energy as a State Function. (a) Work (b) Heat (c) The general case, the first law of Thermodynamics; Applications; Work; Enthalpy, H - a useful new state function; Extensive and intensive properties; The relationship between C_p and C_v ; Measurement of ΔU and ΔH : Calorimetry; Enthalpy change, $\Delta_r H$ of reactions- reaction Enthalpy (a) Standard enthalpy of reactions, (b) Enthalpy changes during transformations, (c) Standard enthalpy of formation, (d) Thermochemical equations (e) Hess's law of constant Heat summation; Enthalpies for different types of reactions. (a) Standard enthalpy of combustion ($\Delta_c H^0$), (b) Enthalpy of atomization ($\Delta_a H^0$), phase transition, sublimation and ionization, (c) Bond Enthalpy ($\Delta_{\text{bond}} H^0$), (d) Enthalpy of solution ($\Delta_{\text{sol}} H^0$) and dilution-lattice enthalpy; Spontaneity. (a) Is decrease in enthalpy a criterion for spontaneity? (b) Entropy and spontaneity, the second law of thermodynamics, (c) Gibbs Energy and spontaneity; Absolute entropy and the third law of thermodynamics.

Unit-7: CHEMICAL EQUILIBRIUM AND ACIDS-BASES: Equilibrium in Physical process; Equilibrium in chemical process - Dynamic Equilibrium; Law of chemical Equilibrium - Law of mass action and Equilibrium constant; Homogeneous Equilibria, Equilibrium constant in gaseous systems. Relationship between K_p and K_c ; Heterogeneous Equilibria; Applications of Equilibrium constant; Relationship between Equilibrium constant K , reaction quotient Q and Gibbs energy G ; Factors affecting Equilibria.-Le-chatlier principle application to industrial synthesis of Ammonia and Sulphur trioxide; Ionic Equilibrium in solutions; Acids, bases and salts- Arrhenius, Bronsted-Lowry and Lewis concepts of acids and bases; Ionisation of Acids and Bases - Ionisation constant of water and its ionic product- P^H scale-ionisation constants of weak acids-ionisation of weak bases-relation between K_a and K_b -Di and poly basic acids and di and poly acidic Bases-Factors affecting acid strength- Common ion effect in the ionization of acids and bases- Hydrolysis of salts and pH of their solutions, Buffer solutions - Solubility Equilibria of sparingly soluble salts. Solubility product constant Common ion effect on solubility of Ionic salts.

Unit-8: HYDROGEN AND ITS COMPOUNDS: Position of hydrogen in the periodic table; Dihydrogen-Occurrence and Isotopes; Preparation and properties of dihydrogen, uses of H_2 , Hydrides: Ionic, covalent, and non-stoichiometric hydrides; Water: Physical properties; structure of water, ice. Chemical properties of water; hard and soft water, Temporary and permanent hardness of water; Hydrogen Peroxide: Preparation, properties, structure, storage and uses. Heavy Water; Hydrogen as a fuel.

Unit-9: THE s-BLOCK ELEMENTS (ALKALI AND ALKALINE EARTH METALS):

Group 1 Elements : Alkali metals; Electronic configurations; Atomic and Ionic radii; Ionization enthalpy; Hydration enthalpy; Physical properties; Chemical properties; Uses; General characteristics of the compounds of the alkali metals: Oxides; Halides; Salts of oxo Acids; Anomalous properties of Lithium: Differences and similarities with other alkali metals, Diagonal relationship; similarities between Lithium and Magnesium; Some important compounds of Sodium: Sodium Chloride, Sodium carbonate, Sodium Hydroxide, Sodium Bicarbonate, Biological importance of Sodium and Potassium.

Group 2 Elements: Alkaline earth elements; Electronic configuration; Ionization enthalpy;

Hydration enthalpy; Physical properties, Chemical properties; Uses; General characteristics of compounds of the Alkaline Earth Metals: Oxides, hydroxides, halides, salts of oxy acids (Carbonates; Sulphates and Nitrates); Anomalous behavior of Beryllium; its diagonal relationship with Aluminum; Some important compounds of calcium: Preparation and uses of Calcium Hydroxide, Quick lime, Calcium Carbonate, Plaster of Paris; Cement, Biological importance of Calcium and Magnesium.

Unit-10: p- BLOCK ELEMENTS GROUP 13 (BORON FAMILY): General introduction – Electronic configuration, atomic radii, Ionization enthalpy, Electro negativity; Physical & Chemical properties Aluminum reactivity towards acids & alkalies, Important trends and anomalous properties of boron; Some important compounds of Boron- Borax, orthoboric acid, Diborane, Uses of boron, aluminum, and their compounds.

Unit-11: p-BLOCK ELEMENTS - GROUP 14 (CARBON FAMILY): General introduction – Electronic configuration, atomic radii, Ionization enthalpy, Electro negativity; Physical & Chemical properties; Important trends and anomalous properties of carbon; Allotropes of carbon; Uses of carbon; Some important compounds of Carbon and Silicon: Carbon Monoxide, Carbon dioxide, Silica, Silicones, Silicates, Zeolites.

Unit-12: Environmental Chemistry: Definition of terms: Air, Water, Soil Pollutions, Environmental Pollution, Atmospheric Pollution, Acid rain, Particulate pollutants, Stratospheric pollution, Water pollution, Soil pollution, Strategies to control Environmental pollution, Green Chemistry.

Unit-13: ORGANIC CHEMISTRY-SOME BASIC PRINCIPLES AND TECHNIQUES AND HYDRO CARBONS: General introduction; Tetravalency of Carbon: shapes of organic compounds; Structural representations of organic compounds; Classification of organic compounds; Nomenclature of organic compounds; Isomerism; Fundamental concepts in organic reaction mechanisms; Fission of covalent bond; Nucleophiles and electrophiles; Electron movements in organic reactions; Electron displacement effects in covalent bonds: inductive effect, resonance, resonance effect, electromeric effect, hyper conjugation; Types of Organic reactions; Methods of purification of Organic compounds, Qualitative elemental analysis of Organic compounds, Quantitative elemental analysis.

Hydrocarbons: Classification of Hydrocarbons; **Alkanes** - Nomenclature, isomerism (structural and conformations of ethane only); Preparation of alkanes; Properties - Physical properties and chemical Reactivity, Substitution reactions – Halogenation, Combustion, Controlled Oxidation, Isomerisation, Aromatization, and reaction with steam, pyrolysis; **Alkenes-** Nomenclature, structure of ethene, Isomerism (structural and geometrical) ; Methods of preparation; Properties- Physical and chemical reactions: Addition of dihydrogen, halogen, water, Sulphuric acid, Hydrogen halides (Mechanism- ionic and peroxide effect, Markovnikov's, anti-Markovnikov's or Kharasch effect). Oxidation, Ozonolysis and Polymerization; **Alkynes** - Nomenclature and isomerism, structure of acetylene. Methods of preparation of acetylene; Physical properties, Chemical reactions- acidic character of alkyne, addition reactions-of hydrogen, Halogen, Hydrogen halides and water. Polymerization; **Aromatic Hydrocarbons:** Nomenclature and isomerism, Structure of benzene, Resonance and aromaticity; Preparation of benzene. Physical properties. Chemical properties: Mechanism of electrophilic substitution. Electrophilic substitution reactions- Nitration, Sulphonation, Halogenation, Friedel- Craft's alkylation and acylation; Directive influence of functional groups in mono substituted benzene, Carcinogenicity and toxicity.

Unit-14: SOLID STATE: General characteristics of solid state; Amorphous and crystalline solids; Classification of crystalline solids based on different binding forces (molecular, ionic, metallic and covalent solids); Probing the structure of solids: X-ray crystallography; Crystal lattices and unit cells. Bravais lattices primitive and centered unit cells; Number of atoms in a unit

cell (primitive, body centered and face centered cubic unit cell); Close packed structures: Close packing in one dimension, in two dimensions and in three dimensions- tetrahedral and octahedral voids- formula of a compound and number of voids filled- locating tetrahedral and octahedral voids; Packing efficiency in simple cubic, bcc and in hcp, ccp lattice; Calculations involving unit cell dimensions-density of the unit cell; Imperfections in solids-types of point defects-stoichiometric and non-stoichiometric defects; Electrical properties-conduction of electricity in metals, semiconductors and insulators- band theory of metals; Magnetic properties

Unit-15: SOLUTIONS: Types of solutions; Expressing concentration of solutions-mass percentage, volume percentage, mass by volume percentage, parts per million, mole fraction, molarity and molality; Solubility: Solubility of a solid in a liquid, solubility of a gas in a liquid, Henry's law; Vapour pressure of liquid solutions: vapour pressure of liquid- liquid solutions. Raoult's law as a special case of Henry's law -vapour pressure of solutions of solids in liquids; Ideal and non-ideal solutions; Colligative properties and determination of molar mass-relative lowering of vapour pressure-elevation of boiling point-depression of freezing point-osmosis and osmotic pressure-reverse osmosis and water purification; Abnormal molar masses-van't Hoff factor

Unit-16: ELECTROCHEMISTRY AND CHEMICAL KINETICS:

Electrochemistry: Electrochemical cells; Galvanic cells: measurement of electrode potentials; Nernst equation- equilibrium constant from Nernst equation- electrochemical cell and Gibbs energy of the cell reaction; Conductance of electrolytic solutions-measurement of the conductivity of ionic solutions-variation of conductivity and molar conductivity with concentration-strong electrolytes and weak electrolytes-applications of Kohlrausch's law; Electrolytic cells and electrolysis: Faraday's laws of electrolysis-products of electrolysis; Batteries: primary batteries and secondary batteries, Fuel cells, Corrosion of metals-Hydrogen economy.

Chemical Kinetics: Rate of a chemical reaction; Factors influencing rate of a reaction: dependence of rate on concentration- rate expression and rate constant-order of a reaction, molecularity of a reaction; Integrated rate equations-zero order reactions-first order reactions- half-life of a reaction; Pseudo first order reactions; Temperature dependence of the rate of a reaction - effect of catalyst; Collision theory of chemical reaction rates

Unit-17: SURFACE CHEMISTRY: Adsorption: Distinction between adsorption and absorption-mechanism of adsorption- types of adsorption- characteristics of physisorption-characteristics of chemisorption- adsorption isotherms- adsorption from solution phase-applications of adsorption;

Catalysis: Catalysts, promoters and poisons-autocatalysis- homogeneous and heterogeneous catalysis-adsorption theory of heterogeneous catalysis- important features of solid catalysts: (a)activity (b)selectivity- shape-selective catalysis by zeolites- enzyme catalysis-characteristics and mechanism- catalysts in industry

Colloids: Classification of colloids: Classification based on physical state of dispersed phase and dispersion medium-classification based on nature of interaction between dispersed phase and dispersion medium- classification based on type of particles of the dispersed phase-multimolecular, macromolecular and associated colloids-cleansing action of soaps-preparation of colloids-purification of colloidal solutions- properties of colloidal solutions: Colligative properties, Tyndal effect, colour, Brownian movement-charge on colloidal particles, electrophoresis; coagulation-precipitation methods-coagulation of lyophilic sols and protection of colloids-Emulsions; Colloids around us-application of colloids.

Unit-18: GENERAL PRINCIPLES OF METALLURGY: Occurrence of metals; Concentration of ores- levigation, magnetic separation, froth floatation leaching; Extraction of crude metal from concentrated ore-conversion to oxide, reduction of oxide to the metal; Thermodynamic principles

of metallurgy-Ellingham diagram-limitations- applications-extraction of iron, copper and zinc from their oxides; Electrochemical principles of metallurgy; Oxidation and reduction; Refining of crude metal-distillation, liquation poling, electrolysis, zone refining and vapour phase refining; Uses of aluminum, copper, zinc and iron

Unit-19: p-BLOCK ELEMENTS: Group-15 Elements: Occurrence-electronic configuration, atomic and ionic radii, ionization enthalpy, electronegativity, physical and chemical properties; Dinitrogen-preparation, properties and uses; Compounds of nitrogen-preparation, properties, and uses of ammonia; Oxides of nitrogen; Preparation and properties of nitric acid; Phosphorous-allotropic forms; Phosphine- preparation and properties; Phosphorous halides; Oxoacids of phosphorous; Phosphorous halides & Oxo acids of phosphorous

Group-16 Elements: Occurrence-electronic configuration, atomic and ionic radii, ionization enthalpy, electron gain enthalpy, electronegativity, physical and chemical properties; Dioxygen-preparation, properties and uses; Simple oxides; Ozone-preparation, properties, structure and uses; Sulphur-allotropic forms; Sulphur dioxide-preparation, properties and uses; Oxoacids of sulphur; Sulphuric acid-industrial process of manufacture, properties and uses.

Group-17 Elements: Occurrence, electronic configuration, atomic and ionic radii, ionization enthalpy, electron gain enthalpy, **electronegativity**, physical and chemical properties; Chlorine-preparation, properties and uses; Hydrogen chloride- preparation, properties and uses; Oxoacids of halogens; Interhalogen compounds-preparation, properties and uses.

Group-18 Elements: Occurrence, electronic configuration, ionization enthalpy, atomic radii, electron gain enthalpy, physical and chemical properties(a) Xenon-fluorine compounds- XeF_2 , XeF_4 and XeF_6 -preparation, hydrolysis and formation of fluoro anions-structures of XeF_2 , XeF_4 and XeF_6 (b) Xenon-oxygen compounds XeO_3 and XeOF_4 -their formation and structures-uses of noble gases.

Unit-20: d AND f BLOCK ELEMENTS & COORDINATION COMPOUNDS: d and f block elements: Position in the periodic table; Electronic configuration of the d-block elements; General properties of the transition elements (d-block) -physical properties, variation in atomic and ionic sizes of transition series, ionization enthalpies, oxidation states, trends in the M^{2+}/M and $\text{M}^{3+}/\text{M}^{2+}$ standard electrode potentials, trends in stability of higher oxidation states, chemical reactivity and E^θ values, magnetic properties, formation of coloured ions, formation of complex compounds, catalytic properties, formation of interstitial compounds, alloy formation; Some important compounds of transition elements-oxides and oxo-anions of metals-uses of potassium dichromate and potassium permanganate-structures of chromate, dichromate, manganate and permanganate ions; Inner transition elements(f-block)-lanthanoids-electronic configuration-atomic and ionic sizes-oxidation states- general characteristics; The Actinoids- electronic configurations, ionic sizes, oxidation states, general characteristics and comparison with lanthanoids; Some applications of d and f block elements.

Coordination compounds: Werner's theory of coordination compounds; Definitions of some terms used in coordination compounds; Nomenclature of coordination compounds-IUPAC nomenclature;

Isomerism in coordination compounds-(a) Stereo isomerism- Geometrical and optical isomerism (b) Structural isomerism- linkage, coordination, ionization and solvate isomerism

Bonding in coordination compounds. (a) Valence bond theory - magnetic properties of coordination compounds-limitations of valence bond theory (b) Crystal field theory (i) Crystal field splitting in octahedral and tetrahedral coordination entities (ii) Colour in coordination compounds- limitations of crystal field theory; Bonding in metal carbonyls; Stability of coordination compounds; Importance and applications of coordination compounds.

Unit-21: POLYMERS: Classification of Polymers -Classification based on source, structure, mode of polymerization, molecular forces and growth polymerization; Types of polymerization reactions- addition polymerization or chain growth polymerization- Ionic polymerization, free radical mechanism-preparation of addition polymers- polythene, Teflon and poly acrylonitrile- condensation polymerization or step growth polymerization-polyamides- preparation of Nylon 6,6 and nylon 6-polyesters- terylene- bakelite, melamine, formaldehyde polymer- copolymerization- Rubber- natural rubber-vulcanisation of rubber-Synthetic rubbers- preparation of neoprene and buna-N; Molecular mass of polymers-number average and weight average molecular masses- poly dispersity index(PDI); Biodegradable polymers- poly β -Hydroxy butyrate-Co β -Hydroxy Valerate (PHBV), Nylon 2-nylon 6; Polymers of commercial importance- poly propene, polystyrene, poly vinyl chloride(PVC), urea- formaldehyde resin, glyptal, bakelite- their monomers, structures and uses

Unit-22: BIOMOLECULES: Carbohydrates-Classification of carbohydrates-Monosaccharides: preparation of glucose from sucrose and starch-Properties and structure of glucose-D, L configurations and (+), (-) configurations of glucose-Structure of fructose; Disaccharides: Sucrose- preparation, structure; Invert sugar- Structures of maltose and lactose- Polysaccharides: Structures of starch, cellulose and glycogen- Importance of carbohydrates; Amino acids: Natural amino acids- classification of amino acids-structures and D and L forms-Zwitterions; Proteins-Structures, classification, fibrous and globular- primary, secondary, tertiary and quaternary structures of proteins- Denaturation of proteins;

Enzymes: Enzymes, mechanism of enzyme action; **Vitamins:** Explanation-names- classification of vitamins - sources of vitamins-deficiency diseases of different types of vitamins; **Nucleic acids:** chemical composition of nucleic acids, structures of nucleic acids, DNA fingerprinting biological functions of nucleic acids;

Hormones: Definition, different types of hormones, their production, biological activity, diseases due to their abnormal activities.

Unit-23: CHEMISTRY IN EVERYDAYLIFE- Drugs and their classification: (a) Classification of drugs on the basis of pharmacological effect (b) Classification of drugs on the basis of drug action (c) Classification of drugs on the basis of chemical structure (d) Classification of drugs on the basis of molecular targets; Drug-Target interaction-Enzymes as drug targets (a) Catalytic action of enzymes (b) Drug-enzyme interaction Receptors as drug targets; Therapeutic action of different classes of drugs: antacids, antihistamines, neurologically active drugs: tranquilizers, analgesics- non- narcotic, narcotic analgesics, antimicrobials-antibiotics, antiseptics and disinfectants-antifertility drugs; Chemicals in food- artificial sweetening agents, food preservatives, antioxidants in food; Cleansing agents-soaps and synthetic detergents

Unit-24:HALO ALKANES AND HALO ARENES: Classification and nomenclature; Nature of C-X bond; Methods of preparation: Alkylhalides and arylhalides-from alcohols, from hydrocarbons (a) by free radical halogenation (b) by electrophilic substitution (c) by replacement of diazonium group (Sandmeyer reaction) (d) by the addition of hydrogen halides and halogens to alkenes-by halogen exchange reactions (Finkelstein reaction, Swarts reaction); Physical properties-melting and boiling points, density and solubility; Chemical reactions: Reactions of haloalkanes (i) Nucleophilic substitution reactions (a) SN^2 mechanism (b) SN^1 mechanism (c) stereo chemical aspects of nucleophilic substitution reactions-optical activity (ii) Elimination reactions (iii) Reaction with metals-Reactions of haloarenes: (i) Nucleophilic substitution (ii) Electrophilic substitution and (iii) Reaction with metals; Polyhalogen compounds: Uses and environmental effects of dichloro methane, trichloromethane, triiodomethane, tetrachloro methane, freons and DDT

Unit-25: ORGANIC COMPOUNDS CONTAINING C, H AND O (ALCOHOLS, PHENOLS, ETHERS, ALDEHYDES, KETONES AND CARBOXYLIC ACIDS):

Alcohols, Phenols and Ethers: Alcohols, phenols and ethers-classification; Nomenclature: (a) Alcohols, (b) phenols and (c) ethers; Structures of hydroxy and ether functional groups; Methods of preparation: Alcohols from alkenes and carbonyl compounds, from Grignard reagents; Phenols from haloarenes, benzene sulphonic acid, diazonium salts, cumene; Physical properties of alcohols and phenols; Chemical reactions of alcohols and phenols (i) Reactions involving cleavage of O-H bond in alcohols-Acidity of alcohols and phenols, esterification (ii) Reactions involving cleavage of C-O bond-reactions with HX, PX_3 , dehydration and oxidation (iii) Reactions of phenols-electrophilic aromatic substitution, Kolbe's reaction, Reimer - Tiemann reaction, reaction with zinc dust, oxidation;

Commercially important alcohols (methanol, ethanol)

Ethers-Methods of preparation: By dehydration of alcohols, Williamson synthesis-Physical properties-Chemical reactions: Cleavage of C-O bond and electrophilic substitution of aromatic ethers (anisole).

Aldehydes and Ketones: Nomenclature and structure of carbonyl group; Preparation of aldehydes and ketones-(1) by oxidation of alcohols (2) by dehydrogenation of alcohols (3) from hydrocarbons -Preparation of aldehydes (1) from acyl chlorides (2) from nitriles and esters (3) from hydrocarbons-Preparation of ketones (1) from acyl chlorides (2) from nitriles (3) from benzene or substituted benzenes; Physical properties of aldehydes and ketones; Chemical reactions of aldehydes and ketones-nucleophilic addition, reduction, oxidation, reactions due to α -Hydrogen and other reactions (Cannizzaro reaction, electrophilic substitution reaction); Uses of aldehydes and ketones.

Carboxylic acids: Nomenclature and structure of carboxyl group; Methods of preparation of carboxylic acids (1) from primary alcohols and aldehydes (2) from alkyl benzenes (3) from nitriles and amides (4) from Grignard reagents (5) from acyl halides and anhydrides (6) from esters; Physical properties; Chemical reactions: (i) Reactions involving cleavage of O-H bond-acidity, reactions with metals and alkalis (ii) Reactions involving cleavage of C-OH bond- formation of anhydride, reactions with PCl_5 , PCl_3 , $SOCl_2$, esterification and reaction with ammonia (iii) Reactions involving-COOH group-reduction, decarboxylation (iv) Substitution reactions in the hydrocarbon part-halogenation and ring substitution; Uses of carboxylic acids.

Unit-26: ORGANIC COMPOUNDS CONTAINING NITROGEN:

Amines: Structure of amines; Classification; Nomenclature; Preparation of amines: reduction of nitro compounds, ammonolysis of alkyl halides, reduction of nitriles, reduction of amides, Gabriel phthalimide synthesis and Hoffmann bromamide degradation reaction; Physical properties; Chemical reactions: basic character of amines, alkylation, acylation, carbyl amine reaction, reaction with nitrous acid, reaction with aryl sulphonyl chloride, electrophilic substitution of aromatic amines (aniline)-bromination, nitration and sulphonation.

DIAZONIUM SALTS- Methods of preparation of diazonium salts (by diazotization); Physical properties; Chemical reactions: Reactions involving displacement of Nitrogen, reactions involving retention of di azo group –coupling reactions; Importance of diazonium salts in synthesis of aromatic compounds

Cyanides and Isocyanides:

Structure and nomenclature of cyanides and isocyanides; Preparation, physical properties and chemical reactions of cyanides and isocyanide

ANNEXURE – II

MODEL QUESTIONS – MATHEMATICS

- 1) If a point (a,a) falls between the lines $|x+y| = 4$
1) $|a| = 2$ 2) $|a| = 3$ 3) $|a| < 2$ 4) $|a| < 3$
- 2) The variance of 30 observations is 3. If each of the observations is multiplied by 3, then the variance of the resulting observations is :
1) 3 2) 9 3) 27 4) 81
- 3) If the sum of two positive numbers is k, then the sum of their squares will be minimum, when the numbers are
1) $k/4, k/4$ 2) $k/3, k/3$ 3) $k/2, k/2$ 4) k,k
- 4) If the lines $4x+3y-1=0$, $x-y+5=0$ and $kx+5y-3=0$ are concurrent then k =
1) 4 2) 5 3) 6 4) 7
- 5) In any ΔABC , $b^2 \sin 2C + c^2 \sin B =$
1) Δ 2) 2Δ 3) 3Δ 4) 4Δ

MODEL QUESTIONS – PHYSICS

1. A particle starts from origin at $t=0$ with a velocity of $10 \hat{i}$ m/s and moves in x-y plane under the action of force which produces a constant acceleration of $(2\hat{i} + 3\hat{j})$ m/s². The y – coordinate in meters of the particle at the instant its x-coordinate is 24m becomes
(1) 12 (2) 6 (3) 18 (4) 3
2. When 0.2 kg of ice at 0°C mixed with 0.5 kg of water at 60°C in a container, the resulting temperature is 10°C . The heat of fusion of ice ($S_{\text{water}} = 4.186 \text{ J/kg/K}$)
(1) $1.31 \times 10^5 \text{ J/kg}$ (2) $2.62 \times 10^5 \text{ J/kg}$
(3) $10.46 \times 10^5 \text{ J/kg}$ (4) $5.23 \times 10^5 \text{ J/kg}$
3. 5 bulbs each of 100 W are connected across 220 V power supply for domestic application. If each unit costs Rs. 4 then the cost per day in Rs. is
(1) 48 (2) 24 (3) 96 (4) 12
4. A solenoid of length 1.0m has a radius of 1cm and is made up of 1000 turns. It carries a current of 2.5 A. The magnitude of the magnetic field inside the solenoid in Teslas
(1) $\pi \times 10^{-3}$ (2) $\pi \times 10^{-4}$ (3) $\pi \times 10^{-6}$ (4) $\pi \times 10^{-5}$

MODEL QUESTIONS – CHEMISTRY

1. Which one of the following has stable electronic configuration?
(1) N (2) C (3) F (4) Al
2. Which one of the following exhibits acidity?
(1) R-OH (2) R-CHO (3) R-X (4) $\text{C}_6\text{H}_5\text{-OH}$
3. Assertion (A): Carbonyl compounds undergo nucleophilic addition reactions.

Reason (R): Carbonyl group is non-polar.

The correct answer is:

- (1) Both (A) and (R) are true and (R) is the correct explanation of(A)
- (2) Both (A) and (R) are true and (R) is not the correct explanation of(A)
- (3) (A) is true but (R) is not true
- (4) (A) is not true but (R) is true

4. Match the following:

LIST I LIST II

- | | |
|--|-----------|
| (A) Packing efficiency in ccp structure | (1) 2 |
| (B) Number of atoms in bcc unit cell | (2) 4 |
| (C) Packing efficiency in simple cubic structure | (3) 52.4% |
| (D) Number of atoms in fcc unit cell | (4) 68.0% |
| | (5) 74% |

The correct answer is:

	(A)	(B)	(C)	(D)
(1)	5	4	3	2
(2)	3	2	1	4
(3)	5	1	3	2
(4)	4	1	2	3

ANNEXURE – III

DEFINITION OF LOCAL / NON – LOCAL STATUS

As per the orders issued by the Government of Andhra Pradesh for 2026-27

ANNEXURE – IV

NORMALIZATION PROCEDURE

Candidates are aware that the AP EAPCET-2026 (MPC and Bi.PC Streams) will be conducted from 12-05-2026 to 18-05-2026 in multiple sessions.

APEAPCET-2026 is being conducted in multiple sessions based on the same syllabus, same pattern for candidates having the same eligibility criteria. A candidate will be eligible to appear only in one session. Since the question paper will be different for each session, there is a possibility that the candidates compare themselves about the variation in the difficulty level of questions. However, it may be noted that utmost care will be taken so that all the papers are of the same standard. Further, it is decided to adopt a normalization process to eliminate any such variations in the difficulty level of various sessions.

What is Normalization?

Normalization, as used in the Indian context, is a process for ensuring the students neither advantaged nor disadvantaged by the difficulty of examinations conducted in multiple sessions. This process is based on a simple formula which has been adopted as recommended by the experts from reputed educational institutions at all India level and Universities. The process is being implemented in other all India / Nationwide entrance tests for admission into undergraduate and graduate professional courses. Normalization process ranks all the candidates across all sessions on a comparative scale. In any normalization process, the marks of the easier session may be reduced marginally and the marks of the harder paper may increase marginally on the global level, depending on the average performance in each session. If there is no much difference in the averages between the two sessions then there won't be much difference in the normalized marks as well. Normalizing marks would justify the candidates while protecting their actual performance.

AP EAPCET marks Normalization Process:

The main aim of the normalization is to justify the candidates who got a difficult paper compared to an easier paper. Hence, the task is to rationalize in the best possible sense and rank the candidates based on the global performance. Various national level examination bodies like JEE (Main), GATE etc. are currently adopting such normalization procedures. Correspondingly, EAPCET committee has deliberated extensively and decided to use the following normalization procedure.

Normalized Marks of the candidate

$$GMS + \frac{Top\ Average\ Global - GMS}{Top\ Average\ Session - SMS} \times (Marks\ Obtained\ by\ Candidate - SMS)$$

where

SMS: (Average + Standard Deviation) of the session in which the candidate belongs to

GMS: (Average + Standard Deviation) of all the candidates across all sessions together

Top Average Session: Average marks of the top 0.1% of the candidates in the session in

which the candidate belongs to

Top Average Global: Average marks of the top 0.1% of all the candidates across all sessions together

Weightage for assigning merit ranks:

75% of AP EAPCET normalized marks and 25% of Intermediate Marks in group subjects to prepare the rank.

Note:

- For Candidates having qualifying marks in AP EAPCET-2026, if after normalization, the marks(s) in any individual subject(s) become negative, then the normalized mark(s) in the respective subject(s) are treated as zero. However, total marks in three subjects are considered as EAPCET marks.
- For the candidates for whom there is no qualifying cut-off in AP EAPCET - 2026, if the marks in all the three subjects after normalization go below zero (negative), the total marks are treated as zero and the rank is assigned. If the tie persists then APEAPCET 2026 normalization marks (though negative are considered for breaking the tie).

Demonstration with a sample data:

The following is based on sample data to explain the normalization process. The data is based on an almost equal number of candidates in all four sessions. The normalization is shown subject wise so that students get the benefit based on subject wise performance rather than the entire paper in a session.

Averages and Standard Deviations in a particular session and averages of top 0.1% candidates of a particular session, Global Average and Standard Deviations of all sessions together, Averages of top 0.1% candidates in all sessions is given in Table

1. Example data of normalized marks is shown in Table 2 to Table 5.

Table 1: Averages and Standard Deviations of sample data

		Maths	Physics	Chemistry
Session1	Avg	27.01245	11.44816	13.56629
	Std_Dev	10.23632	4.135746	5.939418
	Top 0.1% Avg	74.28	37.93	37.7
Session2	Avg	27.23746	11.49711	13.69626
	Std_Dev	10.38974	4.177132	6.005731
	Top 0.1% Avg	74.85	38.03	37.93
Session3	Avg	23.8686	10.25933	13.55555
	Std_Dev	7.717783	3.20095	5.403734
	Top 0.1% Avg	70.05	35.55	39
Session4	Avg	23.95383	10.2931	13.55808
	Std_Dev	7.793973	3.212227	5.460391
	Top 0.1% Avg	70.18	36.4	39.38
<i>All sessions together</i>	<i>Global_Avg</i>	<i>25.52725</i>	<i>10.87743</i>	<i>13.60516</i>
	<i>Global_Std_Dev</i>	<i>9.252138</i>	<i>3.764241</i>	<i>5.718592</i>
	Top 0.1% Global Avg	73.92	37.65	38.74

Table 2: Example of Normalized marks in Session 1:

Candidate	Marks	Maths	Physics	Chemistry	Total
C1	Actual Marks	0	0	0	0
	Normalized Marks	-4.6	-1.407	-1.49	-7.498
C2	Actual Marks	8	3	5	16
	Normalized Marks	3.857	1.682	3.845	9.385
C3	Actual Marks	61	16	25	102
	Normalized Marks	59.89	15.07	25.19	100.1
C4	Actual Marks	76	36	38	150
	Normalized Marks	75.75	35.67	39.06	150.5

Table 3: Example of Normalized marks in Session 2:

Candidate	Marks	Maths	Physics	Chemistry	Total
C1	Actual Marks	1	3	4	8
	Normalized Marks	-3.74	1.595	2.595	0.451
C2	Actual Marks	14	9	2	25
	Normalized Marks	9.932	7.771	0.464	18.17
C3	Actual Marks	48	24	33	105
	Normalized Marks	45.69	23.21	33.49	102.4
C4	Actual Marks	78	38	39	155
	Normalized Marks	77.24	37.62	39.88	154.7

Table 4: Example of Normalized marks in Session 3:

Candidate	Marks	Maths	Physics	Chemistry	Total
C1	Actual Marks	0	0	0	0
	Normalized Marks	2.634	0.622	0.957	4.21
C2	Actual Marks	10	5	1	16
	Normalized Marks	12.81	5.83	1.926	20.6
C3	Actual Marks	50	17	31	98
	Normalized Marks	53.52	18.33	30.99	103
C4	Actual Marks	74	39	38	151
	Normalized Marks	77.94	41.24	37.77	157

Table 5: Example of Normalized marks in Session 4:

Candidate	Marks	Maths	Physics	Chemistry	Total
C1	Actual Marks	4	1	2	7
	Normalized Marks	6.457	1.97	2.935	11.4
C2	Actual Marks	19	7	9	35
	Normalized Marks	21.75	8.018	9.641	39.4
C3	Actual Marks	13	6	16	35
	Normalized Marks	15.63	7.01	16.35	39
C4	Actual Marks	67	9	24	100
	Normalized Marks	70.69	10.03	24.01	105
C5	Actual Marks	57	8	35	100
	Normalized Marks	60.49	9.025	34.55	104
C6	Actual Marks	80	38	40	158
	Normalized Marks	83.94	39.26	39.34	163

ANNEXURE – V

CRITERIA FOR RANKING (AP EAPCET – 2026 “E CATEGORY”)

The candidates who have secured qualifying marks in AP EAPCET-2026 and candidates belonging to the category of Scheduled Caste and Schedule Tribe, for whom qualifying marks have not been prescribed, shall be assigned ranking in the order of merit based on combined score obtained by giving 75% weightage to the marks secured in AP EAPCET-2026 and 25% weightage to the marks secured in the relevant group subjects namely Mathematics, Physics, Chemistry of the qualifying examination.

For the preparation of the merit list, in case of more than one student securing the same combined score obtained as mentioned above, the tie shall be resolved to decide the relative ranking by successively considering the following

- (I) The total marks secured in AP EAPCET-2026
- (II) The Marks secured in mathematics in AP EAPCET-2026
- (III) The marks secured in Physics in AP EAPCET-2026
- (IV) If the tie persists, same rank will be given and the older (based on date of birth) will be given preference over the younger at the time of admissions.

The weightage of marks in the case of candidates belonging to the category of Persons of Indian Origin (PIO) / Overseas Citizen of India (OCI) Card Holders, will be decided by a committee constituted by the competent authority.